

## Electron Dot Diagrams

Each of these compounds is covalently bonded. Please write two structures for each compound. First, do a complete Lewis structure (also called dot diagram), and second, do a structure that uses dashes for bonds.

CH <sub>4</sub>	CHCl <sub>3</sub>	CH <sub>2</sub> Cl <sub>2</sub>	CH <sub>3</sub> Cl	CCl <sub>4</sub>	AlCl <sub>3</sub>	AlCl <sub>4</sub> <sup>1-</sup>
CO <sub>2</sub>	HClO	HCN	H <sup>1-</sup>	NH <sub>3</sub>	NH <sub>4</sub> <sup>+</sup>	PCl <sub>3</sub>
SiF <sub>4</sub>	BH <sub>3</sub>	BH <sub>4</sub> <sup>-</sup>	BF <sub>3</sub>	OH <sup>1-</sup>	C <sub>2</sub> H <sub>6</sub>	C <sub>2</sub> H <sub>5</sub> OH
NH <sub>2</sub> <sup>1-</sup>	H <sub>2</sub> O <sub>2</sub>	CH <sub>3</sub> OH	C <sub>2</sub> H <sub>4</sub>	C <sub>2</sub> H <sub>3</sub> Cl	C <sub>2</sub> H <sub>2</sub> Cl <sub>2</sub>	NF <sub>3</sub>
Cl <sub>2</sub> O	H <sub>2</sub> O	C <sub>2</sub> H <sub>2</sub>	C <sub>2</sub> HCl	C <sub>2</sub> Cl <sub>2</sub>	H <sub>3</sub> O <sup>+</sup>	